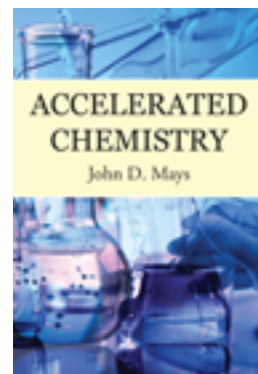


Accelerated Chemistry

Errata

We always strive to make our textbooks as accurate as possible, but sadly, errors are a reality. We very much appreciate friends who report errata that are not included in this document!

Please send new errata to info@centripetalpress.com



Last revised: June 12, 2020

Accelerated Chemistry (2015)

Chapter 1 Exercises

16. Yb: [Xe]6s24f14

Es: [Rn]7s25f11

No: [Rn]7s25f14

28c. 2.91×10^{22} atoms

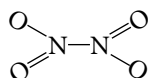
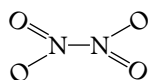
Chapter 2 text

p. 54 The opening of the first paragraph should read, “The first 92 elements...are found in nature. Elements 93–118 have been synthesized in laboratories...”

Chapter 2 Exercises

10. The problem statement should refer to cesium (Cs). Answer: $\text{Mg} < \text{Na} < \text{Ba} < \text{Cs}$

Chapter 3 Exercises



12 o. N₂O₄

22a. The Be—F bond is ionic

Chapter 5

13i. reaction products should be LiI(aq) and K(s)

Chapter 7

14b. 4.20×10^2 kg

Chapter 8

40. the first answer is 3.46 m

Chapter 9

4g. answer is diprotic

21. Add the following note to the answers given in the text: These answers all show the formation of carbonic acid, H_2CO_3 . This acid is unstable and immediately break down to CO_2 and water. Thus, each equation could be shown as: $\dots + \text{CO}_2 + \text{H}_2\text{O}$.

25. The first two sentences of the question should read: According to the activity series of metals (Table 5.2), copper does not react with sulfuric acid. However, if the acid is hot enough and concentrated enough, copper reacts with H_2SO_4 in a single-replacement reaction.

28. basic

Chapter 10 Text

52. 8.50 atm

Chapter 12 Exercises

For exercise 2, the following descriptions should accompany the equations in the answer key.

a. Not a redox reaction.

b. Cl is reduced; it is the oxidizing agent. O is oxidized; it is the reducing agent.

c. S is reduced; it is the oxidizing agent. Br is oxidized; it is the reducing agent.

d. Not a redox reaction.

e. Cl is reduced; it is the oxidizing agent. I is oxidized; it is the reducing agent.

f. N is reduced; it is the oxidizing agent. S is oxidized; it is the reducing agent.

For exercise 7, the following descriptions should accompany the equations in the answer key.

a. oxidizing agent: Fe; reducing agent: S

b. oxidizing agent: Cl; reducing agent: I

c. oxidizing agent: Mn; reducing agent: C

d. oxidizing agent: Cl; reducing agent: O

e. oxidizing agent: N; reducing agent: Al

f. oxidizing agent: Mn; reducing agent: Cl

g. oxidizing agent: N; reducing agent: S

h. oxidizing agent: Mn; reducing agent: Br

14. The second sentence in this question should read: “On your diagram, identify the following: anode, cathode, positive electrode, negative electrode, direction of electron flow, direction of nitrate ion migration in the salt bridge, direction of potassium ion migration in the salt bridge.”

Digital Resources/Resource CD

Exam 2

6. Answer should be 60.052 g/mol

Quiz 5

2. Result should be rounded to hundredths place, giving 24.31.

Fall Semester Exam

- 1d. The compound should be Cl_2O . The answer given is for this compound.
4. Our given solution is correct except for the final result, which should be 1.549×10^{-19} J.
9. The molecular mass of propane used in our solution should be 44.096 g/mol, giving a result of 8.194×10^{25} carbon atoms.
20. Correct answer is $\text{Mg} < \text{Ca} < \text{Sr}^{2+} < \text{Sr} < \text{Ba}^{2+}$

Spring Semester Exam

- 6b. The ionic equation should have $2\text{Ag}^+(\text{aq})$ on both sides (not $2\text{Ag}^{2+}(\text{aq})$)